Review Article

Debaditya Bera#, Rimpa Chatterjee#, and Susanta Banerjee*

Aromatic polyamide nonporous membranes for gas separation application

https://doi.org/10.1515/epoly-2021-0016
received November 12, 2020; accepted January 10, 2021

Abstract: Polymer membrane-based gas separation is a superior economical and energy-efficient separation technique over other conventional separation methods. Over the years, different classes of polymers are investigated for their membrane-based applications. The need to search for new polymers for membrane-based applications has been a continuous research challenge. Aromatic polyamides (PAs), a type of high-performance materials, are known for their high thermal and mechanical stability and excellent film-forming ability. However, their insolubility and processing difficulty impede their growth in membrane-based applications. In this review, we will focus on the PAs that are investigated for membrane-based gas separations applications. We will also address the polymer design principal and its effects on the polymer solubility and its gas separation properties. Accordingly, some of the aromatic PAs developed in the authors’ laboratory that showed significant improvement in the gas separation efficiency and placed them in the 2008 Robeson upper bound are also included in this review. This review will serve as a guide to the future design of PA membranes for gas separations.

Keywords: aromatic polyamide, synthesis, solubility, gas permeability, gas selectivity

1 Introduction

High-performance polymeric materials exhibit several exciting properties, such as high thermal stability, chemical resistance, low flammability, and excellent mechanical properties, that make them useful for advanced technologies. Wholly aromatic polyamides (PAs) are a class of high-performance polymers that find applications in different cutting-edge technologies, particularly in aerospace and military applications. Nevertheless, very high glass transition temperatures of PAs, mostly above their thermal decomposition temperatures, and their low solubility in organic solvents resulted in processing difficulties and restricted many of their applications. Consequently, basic and applied research is directed to enhance their processability and solubility and widen the possibility of their industrial applications.

Separation and purification of gas mixtures are necessary industrial processes in different demanding applications, such as removing carbon dioxide for natural gas “sweetening,” carbon dioxide from greenhouse gas-producing sources, oxygen enrichment from air, etc. Compared to the different traditional separation processes like pressure swing adsorption, cryogenic distillation, etc., gas separation using polymer membranes offers many advantages. The ease of fabrication of polymer membranes, mechanical simplicity, energy efficiency, relatively low cost for small-scale operation, and the overall environment-friendly techniques made them a better choice. The limited thermal and chemical stability, considerable material cost for high output applications, and the inherent trade-off between gas permeability and selectivity of these polymer membranes are the significant issues that need attention. The critical paradigm for the successful application of polymer membrane for cost-effective large-scale industrial applications is to develop polymers with high permeability and good selectivity. Gas molecules permeate through the dense polymer membrane according to the solution diffusion mechanism. According to the model, the permeation of the gas molecules depends on the interaction (solubility) between the
polymer and the individual gas molecules and on the
movement (diffusion) of the gas molecule through the
transient opening of the leap channels (linking between
the free volume elements in the membranes) of adequate
size caused by the segmental mobility of the polymer
chains (10).

In short, the permeability and selectivity of the gas
molecules through the polymer membrane depends on
the interaction between gas molecules and the functional
groups (polar groups) present in the polymer chain and
also on the polymer chain packing density, which con-
trols the fractional-free volume (FFV) in the dense polymer
membrane. However, there is no strict design rule to con-
trol these parameters for obtaining an optimal gas separa-
tion (11). But, generally, it has been observed that pol-
ymers with rigid backbone induce size-based selectivity for
small molecules like CO$_2$ and O$_2$ (kinetic diameter of the
gas molecules, CO$_2$ (3.3 Å), O$_2$ (3.46 Å), N$_2$ (3.64 Å), CH$_4$
(3.8 Å)), and the incorporation of the bulky groups in the
polymer chain, which improve the FFV by limiting the
polymer chain packing, helps to enhanced superior per-
meability and selectivity (12). Different classes of polymers,
such as poly(arylene ether)s (13–16), modified cellulose
(17), and polyimides (PIs) (18–21), have been investigated
for membrane-based gas separation applications. Because
of their certain outstanding qualities like excellent thermal,
mechanical and dimensional stabilities, and low dielectric
constant values, aromatic PIs are known as promising can-
didate for membrane-based gas separation applications
(22). However, the gas separation based on PA membranes
has not gotten proper attention.

Aromatic PAs are considered high-performance mate-
rials because of their versatile chemistry, superior thermal
and mechanical properties, solvent and chemical resist-
ance, high friction resistance, and good film-forming
ability, leading to increased importance in advanced tech-
nologies (1,2,13). Aromatic PAs were first introduced com-
mercially in the early 1960s, in the form of meta-aramid
(Nomex®) produced by DuPont and later para-aramid with
much higher elastic modulus called Kevlar® (Figure 1).
Both the polymers are commercial aromatic PAs of great
economic relevance. PAs draw attention due to their
superior thermal and mechanical properties. PAs are
used for several applications, like protective apparel, air
filtration, thermal and electrical insulation, as advanced
composites in the aerospace and armament industry.
However, difficulties in their processability and solubility
due to strong interchain hydrogen bonding that result in
dense polymer chain packing restricted their use for the
industrial membrane-based gas separations (24).

This review is intended to address in detail the PA
membranes used for gas separation applications. We will
focus on their structural modifications, i.e., strategies
employed to overcome their processability issue and
improve their gas separation properties.

2 PA synthesis

PAs are generally synthesized by condensation polymeri-
ization of an aromatic diamine and an aromatic dicar-
boxylic acid/acid chloride. The synthesis of PAs by dif-
f erent pathways is discussed below.

2.1 Low-temperature polymerization

It is the most common method of preparation of PAs. In
this method, diacid chlorides react with diamines at low
temperatures (from ~10°C to 0°C). This method is gener-
ally preferred when acid chloride of the corresponding
aromatic diacid monomer is easily obtained (Scheme 1).
The solvents used are polar aprotic solvents like N,N-
dimethylacetamide (DMAC), N,N-dimethylformamide
(DMF), and N-methyl-2-pyrrolidone (NMP). Salts, such as CaCl$_2$
and LiCl, are often used as solubility promoters because
the cations interact with the amide groups, diminishing
the interchain hydrogen bonding. During this process,
the main difficulty faces the moisture sensitivity of the
diacid chlorides, which is required in high purity to
obtain the high-molecular-weight polymers (25,26).

Scheme 1: Low-temperature solution polycondensation of diamine
and diacid chloride.

![Figure 1: Structures of the aromatic PAs, Nomex, and Kevlar.](image-url)
2.2 High-temperature polymerization

Yamazaki and Higashi developed this high-temperature polymerization method (27). The condensation reaction occurs directly between diacids and diamines at higher temperatures (from 80°C to 100°C) in the presence of polar aprotic solvents like DMAc, DMF, and NMP. Pyridine and triphenyl phosphate were used as a condensing agent. Also, salts such as CaCl2 and LiCl are used as solubility promoters. Yamazaki and Higashi proposed that the polymerization proceeds via the formation of an intermediate acyloxy N-phosphonium salt of pyridine, as shown in Scheme 2 (27).

Recent efforts have been directed to reduce the polymerization time by applying microwave heating (2 min) instead of the conventional one (4 h) (28). Also, investigations are going on to replace organic solvents like NMP with the “green solvents” like ionic liquids (ILs), which have low vapor pressure, high thermal stability, and high polarity to dissolve the PAs (29).

2.3 Interfacial polycondensation techniques

Interfacial or two-phase polycondensation is one of the promising paths for PA synthesis, and it is a modification of Schotten–Baumann reaction. In this route, monomers are dissolved and reacted in a single solvent path (Scheme 3). Two fast-reacting intermediates dissolved in two immiscible solvents; generally, one of them is water. Usually, the water phase contains an inorganic base for neutralizing acid by-product, and this phase also has the diamine. The other phase includes acid chloride dissolving in organic solvents like hexane, toluene, and dichloromethane. The monomers react at the interface, so the two-phase system was stirred vigorously to obtain a high-molecular-weight product. Low temperature is required to control the side reactions, and the polymers that are unstable at a high temperature can be synthesized (30,31). But the main problem with this technique is the broad molecular weight distribution of polymers, making them unfit for fibers or film fabrication.

2.4 Polycondensation of N-silylated diamine and diacid chlorides

In most of the synthetic efforts, high-molecular-weight PAs were prepared by the activation of diacids. Still, it is one of those few methods which report the activation of diamine for polymer preparation. In this method, diamine can be activated by the reaction with trimethylsilyl chloride. High-molecular-weight polymers were obtained at low temperatures (from −10°C to −5°C) by the polycondensation of N-silylated aromatic diamines and aromatic diacid chlorides (32). This technique involves a subsequent two-step nucleophilic addition–elimination mechanism to react to the acyl-substitution of an acid chloride with an N-silylated amine (Scheme 4) (33). In situ preparation of polymer was reported by Lozano et al., and it is done by the addition of trimethylchlorosilane and diamine to the diacid chloride (21,34).

2.5 Polycondensation of diisocyanates and dicarboxylic acid

The aromatic PAs are prepared by the reaction (at -60°C) of aromatic diisocyanates and dicarboxylic acid (35,36). This method involves the elimination of CO2, and no condensing agent has been used. So various PAs and co-PAs have been synthesized by this method (29,37) (Scheme 5).
2.6 Polycondensation of aromatic diamines, dihalides, and carbon monoxide

The preparation of PAs is also reported by the reaction (at ∼115°C) of aromatic diamines, dihalides, and carbon monoxide in the presence of transition metal catalysts (Pd catalyzed carbonylation coupling of diamines and dihalides; Scheme 6) (21,38). The technique has benefits over the diacid chloride–diamine route due to the ready accessibility of aromatic dibromides and synthesizing of various aramids (38).

3 Structural modifications to improve the processability and gas separation properties

PAs have superior thermal and mechanical properties, making them attractive for different applications, as discussed earlier. However, their processability difficulties hampered their high-end applications. It is difficult to produce thin films out of the PAs or use it for coating because of the PAs’ low solubility in almost all solvents except for concentrated sulfuric acid. The strong and highly directional hydrogen bonding between the PAs’ amide linkages and their aromatic structure results in a high-level intermolecular packing and cohesive energy which is responsible for their inferior processability (1). Besides, the PAs, owing to their structural properties, exhibit extremely high softening and melting temperatures, resulting in their decomposition during or even before melting, ultimately affecting their processability (39). Different strategies have been employed to hinder the chain packing and improve the PAs’ flexibility to enhance their solubility without sacrificing their outstanding thermal and mechanical properties to make them suitable for membrane-based gas separations.

Incorporation of flexible spacers like −O−, −C(CF3)2−, −CH2−, and >C=O into the PA backbone improves their solubility (1). Ferreiro et al. reported the synthesis of several polyisophthalamides by the reaction of three aromatic diamine monomers with four new diacid monomers having oxyethylene groups (40). All the synthesized polymers were soluble in organic aprotic solvents even at room temperature, so processable films were obtained by casting from the respective solutions. The introduction of bulky pendant moiety in the polymer matrix improves their solubility with retention of their thermal properties. Yang et al. reported the synthesis of a series of aromatic PAs from different 2-tert-butylbenzene-substituted diamine and diacid monomers (41). Introducing tert-butyl and aryl ether groups in polymer matrix improves their solubility with retention of their thermal properties. Yagci and Mathias synthesized aromatic PAs by the reaction of 2,5-di-tert-butylbenzene substituted diamine and several diacids and the polymers were soluble in polar aprotic solvents like DMSO, NMP, and DMAc (42).

4 PAs with bulky side substituents and cardo groups

Incorporating the different bulky cardo groups in the backbone of PAs improves its solubility and made it...
possible to process membranes from it for membrane-based gas separation applications. Besides, the cardo groups disrupt the polymer chain packing and enhance the polymer chain’s rigidity, which resulted in improved gas permeability and selectivity. This section will discuss the effect of the specific cardo groups on the gas permeability of PAs.

Polymers bearing fluorinated groups are of great interest to the material chemists because their presence causes noteworthy improvements in individual polymers’ properties compared to their corresponding non-fluorinated analogs. The combination of electronic and steric effects fluorinated groups reduces the interchain interactions and hinders the closer packing of the polymer chains resulting in improved solubility of the polymers (43,44). The C–F bond is a high-energy bond, and the PAs containing fluorine showed excellent thermal properties with enhanced solubility, comparable to those of the conventional aromatic PAs. Incorporation of fluorine as trifluoromethyl (–CF3) (3F) or hexafluoroisopropylidene (–C(F)2–) (6F) groups in the polymer affects several of the polymer properties (43):

- Increases solubility and processability,
- Increases thermal stability,
- Decreases dielectric constant,
- Increases FFV, which resulted in high gas permeability,
- Decreases moisture absorption as well as stability of the polymer membranes.

The greater steric volume of the fluorine relative to hydrogen explained the increase in the polymer’s free volume, which may interfere with efficient chain packing. Moreover, the mutual repulsion between fluorine atoms on different chains may increase the free volume by decreasing the close packing. These excellent balances of properties have made fluorinated polymers very attractive for advanced applications, such as in polymeric membranes for gas separation and pervaporation applications, thermally stable coatings and films, and high-performance structural resins (15,18,44–49). A library of fluorinated monomers (3F group) has been reported in the literature as suitable reactants for fluorinated PA preparation (45,50–56). The bulky 6F and tert-butyl [–C(CH3)3] groups were used extensively to improve the PAs’ processability and gas permeability.

Freeman et al. reported a series of PAs (namely, IP/ SO2, TBI/SO2, and PII/SO2) with different bulky substituents (57). They have prepared the PAs by the reaction of different isophthaloyl chlorides without or with substitution in the fifth position (tert-butyl and phthalimide moiety) and 4,4’-diaminodiphenylsulfone (Figure 2). The reported PAs were amorphous due to the disruption of the polymer chain packing caused by the presence of bulky groups. The PA containing tert-butyl group was soluble in DMAC, whereas the phthalimide moiety containing polymer was soluble in 1:1 mixture of DMAC and DMSO. The PAs containing the bulky substituents showed higher gas permeability than the unsubstituted analog (Table 1). Generally, in a series, the gas permeability increases with an increase in the substituents’ bulkiness. However, the polymers containing the bulkier phthalimide moiety showed lower gas permeability, which has the highest volume of van der Waals (69.40 cm3/mol), than the less bulky tert-butyl group (44.35 cm3/mol) (58). The lower gas permeability of the phthalimide moiety containing PAs is attributed to their polarity induced by enhanced chain packing (higher polarity of the phthalimide moiety).

Interestingly, the >SO2 and the 6F groups significantly affected the gas solubility selectivity of the PIPAs (57,59,60). Considering the concentration of the polar

![Figure 2: Structure of the polyisophthalamides containing different bulky pendant groups (57,59,60).](image-url)
groups in this series of PIPAs, the polymers having 6F moiety showed similar CO2/CH4 solubility selectivity to the many other polymers containing polar carbonyl linkages. Whereas the polymers containing >SO2 variant showed surprisingly substantially lower CO2/CH4 solubility. The lower solubility selectivity of the >SO2 containing PIPAs was attributed to its efficient polymer chain packing responsible for the strong amide–amide linkage interaction. Thereby inhibiting the interaction between the polar carbonyl group of the amide linkage and the CO2 molecule, which is responsible for the observed lower gas solubility selectivity. Also, among all the polymers in this series, PIPAs containing pendant tert-butyl group showed higher gas permeability with a substantial decrease in permeability selectivity than their structural analogs (Table 1). The fact was attributed to the disruption of the polymer chain packing due to the presence of the bulky tert-buty! group, which resulted in higher FFVs (0.110 for TBI/SO2 and 0.153 for TBI/6F) in the polymers than the analogous PIPAs having no bulky pendant groups (0.100 for IP/SO2 and 0.149 for IP/6F). Interestingly the solubility for all the gases like O2, N2, CO2 and CH4 and solubility selectivity of the O2/N2 gas pair were higher for the PAs containing the bulky tert-buty! group (e.g., SCO2 = 2.35, SCH4 = 12.24, SO2 = 1.38, and SO2/SN2 = 2.3 for TBI/6F; and SSCO2 = 9.37, SCH4 = 1.25, SO2 = 0.6, and SO2/SN2 = 1.1 for TBI/SO2) compared to its unsubstituted analogue (e.g., SCO2 = 5.26, SCH4 = 1.17, SO2 = 0.29, and SO2/SN2 = 0.91 for IP/6F; and SSCO2 = 7.48, SCH4 = 1.23, SO2 = 0.56, and SO2/SN2 = 1.0 for IP/SO2). The enhancement of solubility was attributed to the increase in polymer FFV due to incorporating the tert-buty! group. However, in the case of solubility selectivity of CO2/N2, the effect of tert-buty! group is not consistent. Whereas for the PAs having both TBI and 6F moiety (TBI/6F), an increase in CO2/CH4 gas solubility selectivity was observed (SCO2/SCH4 = 5.2) compared to their unsubstituted analogs (SCO2/SCH4 = 4.5 for IP/6F). But for the >SO2 containing polymers, the incorporation of the tert-buty! group (Figure 2) resulted in a decrease in CO2/CH4 gas selectivity (SCO2/SCH4 = 3.7 for TBI/SO2) compared to its unsubstituted analog (SCO2/SCH4 = 6.1 for IP/SO2). This contrasting effect of the tert-buty! group on the CO2/CH4 gas selectivity of PAs was not well understood.

In continuation, Freeman et al. studied the gas transport properties of the PAs (namely, IP/6F, Ph/6F, IP/SO2, and Ph/SO2) synthesized from isophthaloyl chloride derivatives bearing a pendant phenyl substituents at the fifth position and its unsubstituted analogs (60). They have incorporated the 6F group in the PA structures and also prepared their structural analogs containing the >SO2 moiety in place of 6F (Figure 2). In general, the substituted polymers were more permeable but less selective than their unsubstituted analog. Also, high permeability was observed for the fluorinated PAs than the non-fluorinated ones. The effect of the phenyl substitution on the gas permeability of the PAs was more prominent in the >SO2 containing polymers. For instance, in the case of CO2 almost four times increase in gas permeability was observed compared to the nonsubstituted one. Whereas for the phenyl substituted 6F containing PAs, only 2.5 times increase was observed for the CO2 gas. The fact was attributed to the higher significant increase in FFV due to the pendant phenyl substitution for the >SO2 containing polymers (change in FFV = 0.007) than the

<table>
<thead>
<tr>
<th>Polymer FFV</th>
<th>Tg (°C)</th>
<th>Pco2</th>
<th>Po2</th>
<th>α (CO2/CH4)</th>
<th>α (O2/2N2)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>IP/SO2</td>
<td>0.100</td>
<td>323</td>
<td>0.81 0.19 56</td>
<td>6.3</td>
<td></td>
<td>(57)</td>
</tr>
<tr>
<td>TBI/SO2</td>
<td>0.110</td>
<td>337</td>
<td>6.6 2.0 32</td>
<td>5.4</td>
<td></td>
<td></td>
</tr>
<tr>
<td>PII/SO2</td>
<td>0.094</td>
<td>350</td>
<td>2.1 0.44 52</td>
<td>7.3</td>
<td></td>
<td></td>
</tr>
<tr>
<td>IP/6F</td>
<td>0.149</td>
<td>297</td>
<td>9.0 2.4 45</td>
<td>5.7</td>
<td></td>
<td>(59)</td>
</tr>
<tr>
<td>TBI/6F</td>
<td>0.153</td>
<td>315</td>
<td>42.5 11.5 23.7</td>
<td>4.6</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Ph/6F</td>
<td>0.152</td>
<td>311</td>
<td>21 5.5 27</td>
<td>4.6</td>
<td></td>
<td>(60)</td>
</tr>
<tr>
<td>Ph/SO2</td>
<td>0.107</td>
<td>328</td>
<td>3.2 0.75 29</td>
<td>5.8</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

Figure 3: Structure of the PAs having different alky side substituents (24).
butyl groups in addition to the side substituents. The linear analogs 1,4-permeability and gas selectivity for most gas pairs than their chain rigidity. However, instead of a decrease in gas permeability (P) of the polymer chain decreased due to an increase in FFV. The higher FFV values of all these PAs have a significant role in their permselectivity values. The polymers with higher P values than the analogs' polymer structure resulted in higher permselectivity, notably for the less condensable gases.

In another approach, Abajo et al. modified the PAs' structure by the systematic incorporation of different alkyl side substituents, such as methyl, isopropyl, or tert-butyl in the polymer backbone to investigate their effect on the polymer properties (Figure 3) (24). All the PAs having the side substituents showed improved solubility, and membranes were fabricated by solution casting method using DMAC as a solvent. The PAs showed good CO2 and O2 permeability that depends on the side substituents and the linking groups connecting the p-phenylene unit. The connecting groups control the torsional mobility of the polymer chains, i.e., with the change in the connecting group from methylene (PA 4) to isopropylidene linkage (PA 3), the rotational freedom of the polymer chain decreased due to an increase in chain rigidity. However, instead of a decrease in gas permeability as expected to the restricted polymer chain mobility, PA 3 showed the highest gas permeability in the series. The higher FFV (FFV = 1.68) of the connecting isopropylidene group in PA 3 nullified restricted chain mobility of the PA 3 and resulted in high gas permeability. This series of PAs showed relatively high gas permeability (P_co2 = 43 and P_O2 = 9.38 Barrer) with moderate permselectivity (Table 2) due to packing disruptive tert-butyl groups in addition to the side substituents.

Ding and Bikson also investigated the effects of other bulky groups on gas transport properties. They have synthesized a series of PAs reacting 3-(4-carboxyphenyl)-2,3-dihydro-1,1,3-trimethyl-1H-indene-5-carboxylic acid with different aromatic diamines (Figure 4) (61).

All the PAs in this series containing the bulky rigid phenylindane moiety showed high gas permeability (P_co2 = 35.3 and P_O2 = 7.35 Barrer) with high gas separation efficiency (P_co2/P_ch4 = 39 and P_O2/P_N2 = 5.4; Table 3). Interestingly, the PAs prepared from a nonlinear structure, such as 1,3-phenylene (PA 6), exhibited lower gas permeability and gas selectivity for most gas pairs than their linear analog 1,4-phenylene (PA 10). Also, the PAs containing the packing disrupting trifluoromethyl group (PA 8 and PA 9) exhibit higher gas permeability with moderate selectivity than other PAs in this series, which is in line with the observed results discussed earlier.

Generally, the PAs' properties were impacted by the number of bulky groups present in it. Hu et al. systematically varied the number of bulky tert-butyl groups in the PAs and investigated their O2–N2 separation properties (62). They have prepared the PAs from 4,4′-hexafluoroisopropylidene-dibenzoic acid with various aromatic diamines containing a varied number of tert-butyl groups (Figure 5).

They reported that polymer packing density plays a vital role in controlling their gas transport properties. Both gas diffusivity and permeability increased and were proportional to the number of bulky groups in the PAs. The specific volume of the polymers strongly influences the gas transport properties of the PAs. The specific volume and the hole affinity constant (G_H) (related to the free volume) of oxygen of the membranes increases with the increase in the size of the diamine section in the polymer backbone (Table 4). The PA with 2,5 di-tert-butyl benzene (PA 16) showed the highest O2 permeability (P_O2 = 3.81 Barrer) without affecting much O2/N2 selectivity (P_O2/P_N2 = 4.63) and the highest selectivity observed for PA 14 with the highest Tg value in the series.

Aguilar-Vega et al. have prepared several PAs by reacting to a cardo diamine monomer, namely, 4,4′-(9-fluorenylidene) diamine with several substituted and unsubstituted aromatic dicarboxylic acids and studied their gas transport properties (Figure 6) (63).

The PAs derived from the isophthalic acid showed the lowest gas permeability (P_co2 = 5.3 and P_O2 = 1.2 Barrer). In contrast, the analogs’ PA bearing the tert-butyl group in the fifth position showed the highest gas permeability (P_co2 = 37.7 and P_O2 = 8.0 Barrer) as well as high gas diffusion coefficient in this series (Table 5). The high packing density of the former (FFV = 0.134) was due to the meta linked isophthalic acid and was responsible for its low permeation, also reported by other researchers (61). The presence of chain packing.
disrupting tert-butyl group later resulted in high FFV (0.149) in the PA (P19) and was responsible for this observed high gas permeability (Table 5).

Kazama et al. also investigated the gas transport properties of the bis(phenyl)fluorene containing PA obtained from Nippon Steel Corp. of the same structure of PA 17 and compared its properties with other types of polymers like PI, polysulfone, and polycarbonate having similar FFV (64). The incorporation of the cardo fluorene moiety increased the polymer rigidity and FFV. The PA showed moderate CO\textsubscript{2} permeability (7.6 Barrer) with high permselectivity of CO\textsubscript{2} over N\textsubscript{2} (PCO\textsubscript{2}/PN\textsubscript{2} = 37). They observed that the magnitude of gas permeability was mainly due to the combination of gas diffusivity and solubility (effect of the bulky cardo fluorene moiety). In contrast, the magnitude of the permselectivity of CO\textsubscript{2} over N\textsubscript{2} was attributed to the solubility selectivity. However, the reported diffusion coefficient of the gases through the PA was lower compared to the other polymers like PI, polysulfone, polycarbonate, etc., having similar FFV. The intermolecular hydrogen bonding present in the PA hinders the polymer chain motion, leading to the lower apparent diffusion coefficient values compared to the other types of polymers.

The effect of the –CF\textsubscript{3} and tert-butyl groups on the PAs gas transport properties was investigated in detail by Carrera-Figueiras and Aguilar-Vega (65). They prepared a series of PAs from 4,4′-(hexafluoroisopropylidene) diphenylene (HFA) and 5-tert-butylisophthalic acid (TERT) and its unsubstituted analog (ISO; Figure 7). Also, to understand the effect of both these bulky groups they synthesized co-PAs taking a mixture of ISO/TERT. The PA from HFA/TERT showed 2–6 times more permeability than their unsubstituted analog HFA/ISO (Table 6). The presence

Table 3: FFV, T\textsubscript{g}, and gas permeability (in Barrers) and permselectivity values of the PAs having indene moiety measured at 30°C and 1–2 atm upstream pressure (61).

<table>
<thead>
<tr>
<th>Polymer</th>
<th>FFV</th>
<th>T\textsubscript{g} (°C)</th>
<th>P\textsubscript{CO\textsubscript{2}} (Barrer)</th>
<th>P\textsubscript{O\textsubscript{2}} (Barrer)</th>
<th>α (CO\textsubscript{2}/CH\textsubscript{4})</th>
<th>α (O\textsubscript{2}/N\textsubscript{2})</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA 6</td>
<td>0.188</td>
<td>346</td>
<td>5.2</td>
<td>1.16</td>
<td>58</td>
<td>7.6</td>
</tr>
<tr>
<td>PA 7</td>
<td>0.180</td>
<td>302</td>
<td>8.4</td>
<td>1.67</td>
<td>39</td>
<td>6.8</td>
</tr>
<tr>
<td>PA 8</td>
<td>0.185</td>
<td>≈400</td>
<td>35.3</td>
<td>7.35</td>
<td>39</td>
<td>5.4</td>
</tr>
<tr>
<td>PA 9</td>
<td>0.203</td>
<td>≈400</td>
<td>31.2</td>
<td>7.23</td>
<td>35</td>
<td>5.5</td>
</tr>
<tr>
<td>PA 10</td>
<td>0.187</td>
<td>357</td>
<td>8.9</td>
<td>1.67</td>
<td>49</td>
<td>6.8</td>
</tr>
<tr>
<td>PA 11</td>
<td>0.168</td>
<td>323</td>
<td>4.1</td>
<td>0.94</td>
<td>48</td>
<td>8.0</td>
</tr>
<tr>
<td>PA 12</td>
<td>0.174</td>
<td>251</td>
<td>9.6</td>
<td>1.91</td>
<td>29</td>
<td>6.1</td>
</tr>
<tr>
<td>PA 13</td>
<td>0.168</td>
<td>323</td>
<td>2.2</td>
<td>0.53</td>
<td>55</td>
<td>8.8</td>
</tr>
</tbody>
</table>

Figure 4: Structure of the polyamides having indene moiety (61).

Figure 5: Structure of the PAs with varied number of bulky tert-butyl group (62).
of the extra bulky lateral substituent tert-butyl group, in the former, increased its FFV (0.156 for HFA/TERT and 0.143 for HFA/ISO), which resulted in the observed high gas permeability as discussed earlier for other series of PAs. For the co-Ps, the gas permeability increases with the increase in TERT content in the copolymers but gas selectivity reduces. In a similar context, they have prepared another series of PAs containing the benzophenone (DBF) unit in place of HFA (Figure 7) (66). The gas permeability values of these PAs showed a similar trend to the PAs containing the HFA unit (Table 6). In these polymers, the DBF/ISO behaved as a gas barrier, whereas DBF/TERT showed higher permeability compared to DBF/ISO. For instance, in the case of O₂ permeation, the later was about 17 times more permeable than the former due to bulky tert-butyl in it. Another exciting correlation found that the PAs derived from HFA showed much higher gas permeability than analogs’ PAs obtained from DBF. Therefore, it is evident that the 6F group plays a vital role in enhancing gas permeability by manipulating the FFV of the polymers.

In continuation to their effort to modify the PAs for gas separation applications, Aguilar-Vega et al. synthesized rigid block copolyaramides comprising earlier investigated high gas permeable HFA/TERT block and barrier DBF/ISO block (Figure 8) with a different block length of 9:9, 12:12, and 18:18 repeating units on each block (67). The longer block segments showed high permeability and permselectivity (P_{CO₂} = 20.9 and P_{O₂} = 4.7 Barrer, P_{CO₂}/P_{CH₄} = 37.62 and P_{O₂}/P_{N₂} = 6.4) compared to the shorter block segment (P_{CO₂} = 18.5 and P_{O₂} = 4.1 Barrer, P_{CO₂}/P_{CH₄} = 28.68 and P_{O₂}/P_{N₂} = 5.5; Table 7). This was ascribed to the inefficient packing of the copolymer molecules with the increase in the block lengths. A similar observation of gas permeation was also reported for the block copolyaramids when they replaced the DBF/ISO block with the barrier HFA/ISO block (Figure 8) and varied the chain length of different blocks as discussed earlier (68). The presence of additional –CF₃ groups in the barrier block enhanced their permeability more than the copolymers with DBF/ISO unit while having almost similar selectivity. For instance, P_{CO₂} = 70.93 and P_{O₂} = 15.43 Barrer, P_{CO₂}/P_{CH₄} = 33.3 and P_{O₂}/P_{N₂} = 5.04 was observed for the longer block segments in this series of copolymers. Interestingly, this series of co-Ps displayed higher permeability than their parent polymers due to the different building blocks’ inefficient packing. It can be concluded from this work that with proper selection of permeable and barrier block, the block copolymer presents an

**Table 4:** Specific volume, Tₙ, hole affinity constant (Cₙ), gas permeability (in Barrers), and permselectivity values of the PAs with varied number of bulky tert-butyl groups at 1 atm upstream pressure and 35°C (62)

<table>
<thead>
<tr>
<th>Polymer</th>
<th>Specific volume (cm³/g)</th>
<th>Tₙ (°C)</th>
<th>Cₙ</th>
<th>P_{O₂} (Barrer)</th>
<th>α ((O₂/N₂))</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA 14</td>
<td>0.534</td>
<td>262</td>
<td>0.136</td>
<td>1.03</td>
<td>5.02</td>
</tr>
<tr>
<td>PA 15</td>
<td>0.565</td>
<td>215</td>
<td>0.182</td>
<td>2.74</td>
<td>4.92</td>
</tr>
<tr>
<td>PA 16</td>
<td>0.606</td>
<td>220</td>
<td>0.317</td>
<td>3.81</td>
<td>4.63</td>
</tr>
</tbody>
</table>

**Table 5:** FFV, Tₙ, and gas permeability (in Barrers) and permselectivity values of the fluorene containing PAs, measured at 35°C and 2 atm pressure (63)

<table>
<thead>
<tr>
<th>Polymer</th>
<th>FFV</th>
<th>Tₙ (°C)</th>
<th>P_{CO₂} (Barrer)</th>
<th>P_{O₂} (Barrer)</th>
<th>α ((CO₂/CH₄))</th>
<th>α ((O₂/N₂))</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA 17</td>
<td>0.141</td>
<td>319</td>
<td>15.3</td>
<td>3.0</td>
<td>27.9</td>
<td>5.8</td>
</tr>
<tr>
<td>PA 18</td>
<td>0.134</td>
<td>301</td>
<td>5.3</td>
<td>1.2</td>
<td>41.3</td>
<td>6.1</td>
</tr>
<tr>
<td>PA 19</td>
<td>0.149</td>
<td>284</td>
<td>37.7</td>
<td>8.0</td>
<td>24.4</td>
<td>5.7</td>
</tr>
</tbody>
</table>

**Figure 6:** PAs having cardo fluorene derivatives (63).

**Figure 7:** Structure of the PA having different bulky pendant groups (65,66).

**Figure 8:**
advantage on gas permeation over their homopolymer analogs due to the block copolymers’ inefficient packing properties.

Lozano et al. have synthesized two PAs from a rigid diamine monomer with spiro carbon moiety and bulky adamantane in its backbone structure, namely, spiro-(adamantane-2,9′(2′,7′-diamino)-fluorene (SADAF) and isophthaloyl chloride (IPC), and the diacid chloride of 2,2′-bis(4-carboxyphenyl)hexafluoropropylene (6FC) (Figure 9) (69). The adamantane group is relatively free of conformational stress because it consists of three rigid cyclohexane rings with chair conformation, restricting close chain packing. The 6F-SADAF polymer showed an ideal O2 and CO2 permeability coefficients of 3.4 and 14 Barrer, with permselectivity values of 5.37 and 41 for O2/N2 and CO2/CH4, respectively. They also prepared the PAs from the diamines. The PI from SADAF and hexafluoro diphthalic anhydride (6FDA) reported having nearly two times more gas permeabilities with somewhat improvement in gas selectivity. However, no report on the gas permeability values of IP-SADF (P20) is found.

González-Díaz et al. have reported prepared aromatic PAs (PA 22) and several co-PAs, which contain a free hydroxyl pendant group in its repeating unit (70). They further modified PA 22 by post-polymerization technique and introduce different bulky pendant groups in it (Scheme 7) (71).

The polymers showed high solubility due to the presence of –CF3 group and bulky pendant groups in the PA backbone. These modified structures reduce close chain packing, as evident from the FFV values (Table 2). In comparison to their unsubstituted precursor PA 22, the modified polymers possess 13.2–31.4 times higher permeability for CO2 and 3.4–4.3 times more selectivity value for CO2/CH4 (Table 8). This study indicates how nicely the gas permeation properties can be manipulated by simple structural modification of the polymers.

Banerjee et al. is actively working on the development of PA membranes for gas separation applications. His group is designing the diamine monomers to modify the polymer structure to understand the structure–property relation of the polymers for the future development.
of the membrane materials. In their work, flexible ether linkages (–O–) and pendant –CF$_3$ groups were incorporated in the rigid aromatic PAs’ backbone to enhance the polymer solubility and gas permeability. In addition to the above strategies, diamine monomers were synthesized with different cardo groups like cyclohexylidene (72), fluorene (73), and spiro (fluorene-9,9’-xanthene) (74), etc. to obtain the intended properties of the PAs. The synthesized diamines were reacted with the commercially available diacids (Figure 10) to prepare the PAs. Subsequently, the membranes were fabricated thereof and were subjected to gas permeation applications.

The incorporation of the cardo groups in addition to the –CF$_3$ moiety and the flexible ether linkage helps to improve the solubility of the PAs without sacrificing its thermal ($T_{ds} > 200\,^\circ C$) and mechanical properties. All the polymers were amorphous in nature, as a broad amorphous halo were observed in their corresponding X-ray diffractograms (72–74). The amorphous nature is due to the presence of the above bulky structural moieties, which disrupts the π–π stacking of amide and phenyl rings of the aromatic PAs and hinders their chain packing.

All the PAs were readily soluble in a common organic solvent like THF, NMP, DMAc, DMF, and pyridine. PA membranes were prepared by casting from their corresponding DMAc solution showed a good set of mechanical properties. The gas transport properties of the PA membranes were investigated at 35°C and 3.5 Barrers pressure. The gas permeability and selectivity values are given in Table 9. Among all these PAs, polymers containing bulky cardo bis(phenylphenyl)fluorene moiety showed the highest permeability ($P$ in Barrers) for different gases ($P_{CO_2} = 67.42, P_{O_2} = 15.00$) and highest permselectivity

Table 7: Gas permeability (in Barrers) and permselectivity values of the copolyaramids having bulky groups at 2 atm upstream pressure and 35°C

<table>
<thead>
<tr>
<th>Polymer</th>
<th>$P_{CO_2}$ (Barrer)</th>
<th>$P_{O_2}$ (Barrer)</th>
<th>$\alpha$ (CO$_2$/CH$_4$)</th>
<th>$\alpha$ (O$_2$/N$_2$)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>DBF/ISO</td>
<td>0.32</td>
<td>0.069</td>
<td>—</td>
<td>—</td>
<td>(66)</td>
</tr>
<tr>
<td>HFA/TERT-b-DBF/ISO 18</td>
<td>20.9</td>
<td>4.7</td>
<td>37.42</td>
<td>6.4</td>
<td>(67)</td>
</tr>
<tr>
<td>HFA/TERT-b-DBF/ISO 12</td>
<td>16.6</td>
<td>3.3</td>
<td>31.18</td>
<td>7.4</td>
<td></td>
</tr>
<tr>
<td>HFA/TERT-b-DBF/ISO 9</td>
<td>18.5</td>
<td>4.1</td>
<td>28.68</td>
<td>5.5</td>
<td></td>
</tr>
<tr>
<td>HFA/TERT</td>
<td>26.82</td>
<td>5.88</td>
<td>31.18</td>
<td>5.07</td>
<td>(68)</td>
</tr>
<tr>
<td>HFA/TERT-b-HFA/ISO 18</td>
<td>70.93</td>
<td>15.43</td>
<td>33.30</td>
<td>5.04</td>
<td></td>
</tr>
<tr>
<td>HFA/TERT-b-HFA/ISO 12</td>
<td>37.94</td>
<td>8.59</td>
<td>34.80</td>
<td>5.33</td>
<td></td>
</tr>
<tr>
<td>HFA/TERT-b-HFA/ISO 9</td>
<td>40.09</td>
<td>8.27</td>
<td>39.69</td>
<td>5.10</td>
<td></td>
</tr>
<tr>
<td>HFA/ISO</td>
<td>7.60</td>
<td>1.68</td>
<td>42.22</td>
<td>6.00</td>
<td></td>
</tr>
</tbody>
</table>

Figure 9: PAs derived from SADAF (69).

Scheme 7: Scheme for the post modification of the PAs (70).
selectivity for both CO2/CH4 (88.37) and O2/N2 (10.84) gas pairs compared to the other PAs containing cardo bis (phenyl)fluorene and cyclohexyldiene moiety. The order of gas permeability through the PAs containing these bulky cardio groups followed the increasing order of FFV of the PAs, i.e., bis(phenylphenyl)fluorene > bis(phenyl)fluorene > spiro (fluorene-9,9'-xanthene) > cyclohexyldiene. The observation clearly states that in these PAs, the disruption of the polymers’ chain packing due to the steric hindrance of the cardio group plays an important role in controlling their gas transport behavior. Significantly, the higher gas permeability in bis(phenylphenyl)fluorene was attributed to the steric hindrance caused by the extra two phenyl groups, responsible for the reduction in H-bonding and an increase in the effective FFV.

Banerjee et al. took another approach and varied the number of aromatic rings, the pendant group, etc., between the diether linkages (Figure 11) (75–78). They have introduced bulky pendant groups like tert-butyl and adamantane in the phenyl ring and investigated the effect of the different structural modification on the PAs’ properties in detail.

The prepared PAs were amorphous and showed good solubility in different organic solvents, and transparent and flexible membranes were obtained for all PAs from their corresponding DMAc solutions. The incorporation of the tert-butyl groups and adamantyl moiety as the pendant group in polymer backbone improves their gas permeation properties (Table 9) without negatively affecting their properties. For instance, the PAs’ membranes containing the di-tert-butyl groups (PA G-i to G-v-iii and G-v) exhibited high mechanical and thermal stability (tensile strength 2.02 bar upstream pressure (Ti) and 35°C-FiV, and for PA G with extra two tert-butyl moieties. The properties make these PA membranes interesting to investigate for gas separation applications. The PAs containing di-tert-butyl pendant groups exhibited the highest permeability with relatively low gas selectivity (PCO2 = 173 Barrer and PO2 = 44 Barrer, PCO2/PCH4 = 24.03, PO2/PN2 = 7.33 for PA G-i) for different gases. Interestingly these PAs showed higher permeability (Table 9) for bigger penetrant CH4 (kinetic diameter ~3.8 Å) over the smaller N2 (kinetic diameter ~3.64 Å). The fact can be attributed to the higher affinity between the nonpolar CH4 molecule and the alkyl tert-butyl moieties (“like dissolves like”) presence in the PAs (79). The fact was well supported by the observed higher solubility of the CH4 (for PA G-i, SCH4 = 0.80 × 10−2 and for PA G-v SCH4 = 0.56 × 10−2 cm3/cm3 cm Hg) over the N2 (for PA G-I, SN2 = 0.37 × 10−2 and for PA G-v SN2 = 0.28 × 10−2 cm3/cm3 cm Hg). In general, the PAs’ PA G-i to G-iii and G-v with extra two tert-butyl groups showed extremely high permeability with comparable gas selectivity than their analogs’ one (Table 10). For instance, a 12-fold increase in PCO2 and 8-fold increase in PO2 was

![Figure 10](image-url)  
**Figure 10:** Different bulky cardio groups incorporated to modify the polyamides structures (72–74).
observed for the PA G-\textit{v} compared to its analogs' one, PA E-\textit{v}. The fact can be attributed to the reduction in the polymer chain packing and enhancement in the polymer FFV due to packing disruptive bulky tert-butyl groups. The bulky tert-butyl hindered the polymer chains proximity, resulting in lowering H-bondings and related interactions between the polymer chain that enhanced the membrane FFV and gas permeability (Figure 12). Also, the tert-butyl groups increase the polymer chains’ rigidity and restrict their segmental movements, which also help maintain higher or moderate gas selectivity. This effect of the tert-butyl groups is not limited to only the PAs but also observed for other polymer membranes like polyacrylates (80), PIs (81), etc.

<table>
<thead>
<tr>
<th>Polymer</th>
<th>FFV</th>
<th>( T_\text{g} ) (°C)</th>
<th>( P_{\text{CO}_2} )</th>
<th>( P_{\text{O}_2} )</th>
<th>( \alpha \left( \text{CO}_2/\text{CH}_4 \right) )</th>
<th>( \alpha \left( \text{O}_2/\text{N}_2 \right) )</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA A-\textit{i}</td>
<td>0.146</td>
<td>258</td>
<td>30.40</td>
<td>6.30</td>
<td>33.78</td>
<td>6.30</td>
<td>(72)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA A-\textit{ii}</td>
<td>0.159</td>
<td>256</td>
<td>27.50</td>
<td>6.20</td>
<td>34.37</td>
<td>6.89</td>
<td></td>
</tr>
<tr>
<td>PA A-\textit{iii}</td>
<td>0.132</td>
<td>299</td>
<td>9.80</td>
<td>2.80</td>
<td>32.67</td>
<td>5.60</td>
<td></td>
</tr>
<tr>
<td>PA A-\textit{iv}</td>
<td>0.060</td>
<td>n.d.</td>
<td>7.00</td>
<td>1.90</td>
<td>53.84</td>
<td>7.60</td>
<td></td>
</tr>
<tr>
<td>PA A-\textit{v}</td>
<td>0.044</td>
<td>n.d.</td>
<td>4.50</td>
<td>1.60</td>
<td>64.29</td>
<td>10.67</td>
<td></td>
</tr>
<tr>
<td>PA B-\textit{i}</td>
<td>0.161</td>
<td>273</td>
<td>52.00</td>
<td>12.15</td>
<td>60.47</td>
<td>10.21</td>
<td>(73)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA B-\textit{ii}</td>
<td>0.192</td>
<td>297</td>
<td>60.23</td>
<td>12.92</td>
<td>47.05</td>
<td>8.85</td>
<td></td>
</tr>
<tr>
<td>PA B-\textit{iii}</td>
<td>0.179</td>
<td>285</td>
<td>21.53</td>
<td>5.70</td>
<td>43.06</td>
<td>8.14</td>
<td></td>
</tr>
<tr>
<td>PA B-\textit{iv}</td>
<td>0.153</td>
<td>294</td>
<td>19.76</td>
<td>4.93</td>
<td>61.75</td>
<td>8.36</td>
<td></td>
</tr>
<tr>
<td>PA B-\textit{v}</td>
<td>0.144</td>
<td>299</td>
<td>15.92</td>
<td>4.18</td>
<td>54.90</td>
<td>8.20</td>
<td></td>
</tr>
<tr>
<td>PA C-\textit{i}</td>
<td>0.174</td>
<td>274</td>
<td>60.32</td>
<td>13.55</td>
<td>60.32</td>
<td>10.84</td>
<td>(73)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA C-\textit{ii}</td>
<td>0.207</td>
<td>274</td>
<td>67.42</td>
<td>15.00</td>
<td>50.69</td>
<td>8.82</td>
<td></td>
</tr>
<tr>
<td>PA C-\textit{iii}</td>
<td>0.180</td>
<td>272</td>
<td>40.38</td>
<td>9.05</td>
<td>43.06</td>
<td>8.14</td>
<td></td>
</tr>
<tr>
<td>PA C-\textit{iv}</td>
<td>0.160</td>
<td>273</td>
<td>38.00</td>
<td>7.37</td>
<td>88.37</td>
<td>9.21</td>
<td></td>
</tr>
<tr>
<td>PA C-\textit{v}</td>
<td>0.145</td>
<td>278</td>
<td>32.00</td>
<td>6.22</td>
<td>80.00</td>
<td>9.01</td>
<td></td>
</tr>
<tr>
<td>PA D-\textit{i}</td>
<td>0.176</td>
<td>283</td>
<td>48.00</td>
<td>14.00</td>
<td>32.00</td>
<td>7.45</td>
<td>(74)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA D-\textit{ii}</td>
<td>0.192</td>
<td>289</td>
<td>45.00</td>
<td>11.70</td>
<td>35.43</td>
<td>7.80</td>
<td></td>
</tr>
<tr>
<td>PA D-\textit{iii}</td>
<td>0.174</td>
<td>273</td>
<td>14.00</td>
<td>5.10</td>
<td>21.88</td>
<td>8.50</td>
<td></td>
</tr>
<tr>
<td>PA D-\textit{iv}</td>
<td>0.162</td>
<td>278</td>
<td>19.00</td>
<td>7.70</td>
<td>27.14</td>
<td>8.85</td>
<td></td>
</tr>
<tr>
<td>PA D-\textit{v}</td>
<td>0.152</td>
<td>288</td>
<td>14.50</td>
<td>3.80</td>
<td>29.00</td>
<td>7.04</td>
<td></td>
</tr>
<tr>
<td>PA E-\textit{i}</td>
<td>0.133</td>
<td>275</td>
<td>55.41</td>
<td>14.09</td>
<td>24.20</td>
<td>5.92</td>
<td>(75)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA E-\textit{ii}</td>
<td>0.139</td>
<td>281</td>
<td>40.38</td>
<td>9.05</td>
<td>43.06</td>
<td>8.14</td>
<td></td>
</tr>
<tr>
<td>PA E-\textit{iii}</td>
<td>0.150</td>
<td>243</td>
<td>13.01</td>
<td>3.27</td>
<td>32.53</td>
<td>6.17</td>
<td></td>
</tr>
<tr>
<td>PA E-\textit{v}</td>
<td>0.125</td>
<td>290</td>
<td>5.72</td>
<td>2.14</td>
<td>12.43</td>
<td>4.37</td>
<td></td>
</tr>
<tr>
<td>PA F-\textit{i}</td>
<td>0.248</td>
<td>278</td>
<td>42.60</td>
<td>10.00</td>
<td>23.67</td>
<td>4.76</td>
<td>(76)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA F-\textit{ii}</td>
<td>0.257</td>
<td>267</td>
<td>28.70</td>
<td>8.50</td>
<td>26.09</td>
<td>7.08</td>
<td></td>
</tr>
<tr>
<td>PA F-\textit{iii}</td>
<td>0.262</td>
<td>252</td>
<td>14.80</td>
<td>3.70</td>
<td>24.67</td>
<td>5.29</td>
<td></td>
</tr>
<tr>
<td>PA F-\textit{iv}</td>
<td>0.224</td>
<td>n.d.</td>
<td>12.10</td>
<td>3.40</td>
<td>30.25</td>
<td>6.80</td>
<td></td>
</tr>
<tr>
<td>PA F-\textit{v}</td>
<td>0.207</td>
<td>n.d.</td>
<td>8.40</td>
<td>2.30</td>
<td>28.00</td>
<td>5.75</td>
<td></td>
</tr>
<tr>
<td>PA G-\textit{i}</td>
<td>0.190</td>
<td>284</td>
<td>173.00</td>
<td>44.00</td>
<td>24.03</td>
<td>7.33</td>
<td>(77)&lt;sup&gt;a&lt;/sup&gt;</td>
</tr>
<tr>
<td>PA G-\textit{ii}</td>
<td>0.171</td>
<td>276</td>
<td>157.00</td>
<td>40.00</td>
<td>30.78</td>
<td>7.69</td>
<td></td>
</tr>
<tr>
<td>PA G-\textit{iii}</td>
<td>0.153</td>
<td>275</td>
<td>80.00</td>
<td>20.10</td>
<td>26.67</td>
<td>7.44</td>
<td></td>
</tr>
<tr>
<td>PA G-\textit{v}</td>
<td>0.145</td>
<td>286</td>
<td>70.00</td>
<td>17.80</td>
<td>25.93</td>
<td>6.85</td>
<td></td>
</tr>
<tr>
<td>PA H-\textit{i}</td>
<td>0.20</td>
<td>238</td>
<td>34.0</td>
<td>9.5</td>
<td>37.8</td>
<td>7.3</td>
<td>(78)</td>
</tr>
<tr>
<td>PA H-\textit{ii}</td>
<td>0.19</td>
<td>235</td>
<td>29.2</td>
<td>8.2</td>
<td>41.7</td>
<td>8.2</td>
<td></td>
</tr>
<tr>
<td>PA H-\textit{iii}</td>
<td>0.12</td>
<td>219</td>
<td>13.0</td>
<td>3.5</td>
<td>32.5</td>
<td>5.8</td>
<td></td>
</tr>
<tr>
<td>PA I-\textit{i}</td>
<td>0.293</td>
<td>274</td>
<td>164.7</td>
<td>59.7</td>
<td>39.6</td>
<td>5.4</td>
<td>(83)</td>
</tr>
<tr>
<td>PA I-\textit{iii}</td>
<td>0.253</td>
<td>267</td>
<td>81.1</td>
<td>29.7</td>
<td>39.8</td>
<td>5.3</td>
<td></td>
</tr>
<tr>
<td>PA I-\textit{iv}</td>
<td>0.255</td>
<td>272</td>
<td>87.6</td>
<td>34.8</td>
<td>33.1</td>
<td>5.2</td>
<td></td>
</tr>
<tr>
<td>PA I-\textit{v}</td>
<td>0.232</td>
<td>273</td>
<td>82.0</td>
<td>27.8</td>
<td>48.7</td>
<td>5.8</td>
<td></td>
</tr>
<tr>
<td>PA I-\textit{vi}</td>
<td>0.264</td>
<td>264</td>
<td>103.4</td>
<td>45.5</td>
<td>37.7</td>
<td>5.3</td>
<td></td>
</tr>
</tbody>
</table>

<sup>a</sup> FFV values were calculated using the equation FFV = \((V_s - V_o)/V_s\). \(V_o\) was computed from the van der Waals’ volumes (\(V_w\)) of the various groups in the polymer structure provided by Park and Paul (85) using the equation \(V_o = 1.3V_w\), and the specific volume of the polymers (\(V_s\)) was calculated from the density values of the respective polymers and the molar mass of the repeat units.
Also, the PAs having adamantyl moiety exhibited a good combination of permeability and selectivity ($P_{CO_2} = 34$ Barrer and $P_{O_2} = 9.5$ Barrer, $P_{CO_2}/P_{CH_4} = 37.8$, $P_{O_2}/P_{N_2} = 7.3$ for PA H-i. In general, the PAs synthesized from the 4,4’(hexafluoroisopropylidene)bis(benzoic acid) have the best combination of permeability and selectivity (e.g., $P_{CO_2} = 157$ Barrer and $P_{O_2} = 40$ Barrer, $P_{CO_2}/P_{CH_4} = 30.78$, $P_{O_2}/P_{N_2} = 7.69$ for PA G-ii) due to disruption of polymer packing and the restricted torsional motion of phenyl rings around the 6F linkage. Also, for all these series of polymers, the PAs synthesized from naphthalene-2,6-dicarboxylic acid showed relatively lower permeability, which is attributed to the presence of rigid compact naphthalene unit responsible for the lowest FFV and observed gas permeability.

![Figure 11](image1.png)

Table 10: FFV, $T_g$, and gas permeability (in Barrers) and permselectivity values of the TPA containing PAs measured at 35°C and 3.5 bar pressure

<table>
<thead>
<tr>
<th>Polymer</th>
<th>FFV$^{exp}$</th>
<th>FFV$^{sim}$</th>
<th>$T_g$ (°C)</th>
<th>$P_{CO_2}$</th>
<th>$P_{O_2}$</th>
<th>$\alpha$ (CO$2$/CH$4$)</th>
<th>$\alpha$ (O$2$/N$2$)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA J-i</td>
<td>0.143</td>
<td>0.139</td>
<td>303</td>
<td>53.5</td>
<td>13.0</td>
<td>48.64</td>
<td>10.0</td>
<td>(92)</td>
</tr>
<tr>
<td>PA J-ii</td>
<td>0.159</td>
<td>0.163</td>
<td>309</td>
<td>61.5</td>
<td>14.2</td>
<td>51.25</td>
<td>10.14</td>
<td></td>
</tr>
<tr>
<td>PA J-iii</td>
<td>0.088</td>
<td>0.088</td>
<td>299</td>
<td>14.5</td>
<td>3.8</td>
<td>48.33</td>
<td>9.50</td>
<td></td>
</tr>
<tr>
<td>PA J-iv</td>
<td>0.095</td>
<td>0.095</td>
<td>328</td>
<td>23.5</td>
<td>5.0</td>
<td>47.00</td>
<td>8.34</td>
<td></td>
</tr>
<tr>
<td>PA J-v</td>
<td>0.061</td>
<td>0.063</td>
<td>330</td>
<td>12.6</td>
<td>3.5</td>
<td>63.00</td>
<td>11.67</td>
<td></td>
</tr>
<tr>
<td>PA K-i</td>
<td>0.160</td>
<td>0.228</td>
<td>263</td>
<td>86.0</td>
<td>20.2</td>
<td>29.65</td>
<td>8.08</td>
<td>(93)</td>
</tr>
<tr>
<td>PA K-ii</td>
<td>0.188</td>
<td>0.250</td>
<td>271</td>
<td>119.0</td>
<td>29.0</td>
<td>37.19</td>
<td>9.67</td>
<td></td>
</tr>
<tr>
<td>PA K-iii</td>
<td>0.105</td>
<td>0.194</td>
<td>258</td>
<td>51.0</td>
<td>12.8</td>
<td>24.28</td>
<td>7.53</td>
<td></td>
</tr>
<tr>
<td>PA K-iv</td>
<td>0.122</td>
<td>0.202</td>
<td>281</td>
<td>69.7</td>
<td>17.0</td>
<td>30.30</td>
<td>8.5</td>
<td></td>
</tr>
<tr>
<td>PA K-v</td>
<td>0.089</td>
<td>0.166</td>
<td>284</td>
<td>41.5</td>
<td>10.8</td>
<td>24.41</td>
<td>7.71</td>
<td></td>
</tr>
<tr>
<td>PA L-ii</td>
<td>0.315</td>
<td>0.340</td>
<td>291</td>
<td>141</td>
<td>33.4</td>
<td>44.33</td>
<td>8.47</td>
<td>(94)</td>
</tr>
<tr>
<td>PA L-iv</td>
<td>0.243</td>
<td>0.270</td>
<td>273</td>
<td>83</td>
<td>19.8</td>
<td>37.89</td>
<td>6.87</td>
<td></td>
</tr>
<tr>
<td>PA L-v</td>
<td>0.229</td>
<td>0.210</td>
<td>271</td>
<td>78</td>
<td>18.1</td>
<td>43.82</td>
<td>8.22</td>
<td></td>
</tr>
<tr>
<td>PA L-vi</td>
<td>0.266</td>
<td>0.290</td>
<td>283</td>
<td>97</td>
<td>23.6</td>
<td>36.74</td>
<td>7.46</td>
<td></td>
</tr>
</tbody>
</table>

![Figure 12](image2.png)

Figure 12: General cartoon representation on the effect of the bulky groups (tert-butyl) on the polymer chain packing and FFV. (a) Compact polymer chain packing in case of polymer without bulky groups. (b) Disruptive polymer packing and higher free space for the polymers having the bulky groups.
Introduction of 9,10-dihydro-9-oxa-10-phosphapentaphenanthrene 10-oxide (DOPO) into the polymer structure is beneficial in many respect. This unit has a typical spatial conformation free of conformational stress and a bulky structure that hinders close packing between polymer chains and effectively used in polymer to improve their solubility and thermal stability. This phosphorus-containing moiety also improves flame retardancy (82). To take the advantages of DOPO unit, Banerjee et al. has prepared a diamine monomer containing this unit and finally designed a series of PAs (Figure 13) (83). Introduction of bulky DOPO unit restricts efficient packing of polymer chains, chain–chain interactions decreases, and increases solubility (84) and permeability like many other PAs with bulky groups (77). The PAs showed high thermal stability, and \( T_g \) values were in the range of 265–275°C. The polymers showed very high gas permeability \( (P_{CO_2} = 164.7 \text{ and } P_{O_2} = 59.7 \text{ Barrer}) \) and maintained a high selectivity. Gas permeability through the polymer membrane depends on the kinetic diameter of the penetrator gas molecules, packing density (FFV), and mobility of the polymer chains. PA with tert-butyl group and DOPO unit, i.e., PA I-1 possesses the highest FFV and the most increased permeability in the series. The high glass transition temperature of the PAs affects their selectivity value by restricting their intersegmental motion.

5 PAs containing triphenylamine (TPA) moiety

Polymers with TPA core exhibits several unique properties, e.g., electrochromic properties (86) and hole-transporting properties (87). The propeller-shaped TPA structure is also known for packing disruptive group that inhibits the polymer chain packing and draws the attention for developing polymer membranes for gas permeation applications (88–91). Liou et al. used different triphenyl amine-containing PIs for membrane-based gas separation applications (88,89). In addition to the above properties, the propeller-shaped structure improves the solubility of the PAs (86). Accordingly, Banerjee et al. have designed and prepared different TPA core containing diamines that bear other bulky pendant groups like adamantane (92), tri-tert-butyl phenol (93), and trityl (94). They have prepared several PAs using these diamines and investigated the gas permeation properties of the synthesized PAs (Figures 14 and 15).

The PAs with the TPA core showed improved solubility in common organic solvents like NMP, DMAC, DMF, pyridine, and DMSO despite the absence of any fluorinated amines’ solubility improving ether linkage as was in the case for the another PAs discussed in the above sections. The improvement was due to the propeller-shaped structure of the TPA, which inhibits the polymer chain packing. Also, the bulky pendant groups improve the steric hindrance between the polymer chains. The membranes from these polymers were subjected to gas permeability studies for different gases like \( CO_2, O_2, N_2, CH_4, \) etc. (Table 10). The PAs containing the tri tert-butyl phenyl pendant group showed higher permeability for all the gases \( (P_{CO_2} = 119, P_{O_2} = 29 \text{ for PA K-ii}) \) compared to analogs' PAs containing adamantane moiety \( (P_{CO_2} = 61, P_{O_2} = 14 \text{ for PA J-ii}) \). The higher permeability of the former is due to their higher FFV that arises from the three tert-butyl substitutions in the pendant phenyl group. In contrast, the adamantane moiety containing PAs showed high gas selectivity (e.g., \( P_{CO_2}/P_{CH_4} = 63, P_{O_2}/P_{N_2} = 11.67 \) for PA J-v). This is attributed to their more compact chain packing, and high rigidity, which is evident from their higher glass transition temperature (330°C for PA J-v) compared to the tert-butyl phenyl containing analogs (e.g., \( P_{CO_2}/P_{CH_4} = 24.4, P_{O_2}/P_{N_2} = 7.7 \) for PA H-v, \( T_g = 284°C \) PA K-v).

The molecular dynamics’ simulations allowed us to analyze and correlate the free volume distribution (FVD) of elements in these PAs with gas permeability and selectivity. The FVD and sizes of the free volume elements (Figure 16) influenced the polymer membranes gas permeability, diffusivity, and permeselectivity. The FVD largely controls the permeability of small- and medium-sized gas molecules. The FVDs for PAs were following the bulkiness of the diacid moiety in the PAs (Figure 14). Higher free volume elements bigger than 60 \( \AA^3 \) in PA k-i and ii are due to more tert-butyl groups and the 6F moiety, respectively. The diffusion coefficients for all the gases were in concurrence with the more significant free volume elements in the PAs (93).

In continuation to our work, Banerjee et al. incorporated a bulky trityl group into the TPA core-based diamine and used it to synthesize the PAs (Figure 15) (94). The PAs showed high gas permeability with good permeselectivity. As discussed earlier, in the case of this series of PAs, the gas permeability is controlled by their high FFV. The presence of a bulky pendant trityl group is responsible for such high FFVs (0.315 for PA L-ii). The trityl group makes the polymer chain labile, which helps to create micropore in the polymer, responsible for high
FV elements in this series of PAs. The PA L-ii, containing the 6F group, showed highest gas permeability and selectivity in this series \( (P_{CO_2} = 141, P_{O_2} = 33.4, P_{CO_2}/P_{CH_4} = 44.33, P_{O_2}/P_N_2 = 8.47) \), whereas PA L-v having the naphthalene moiety showed the lowest gas permeability \( (P_{CO_2} = 78, P_{O_2} = 18.1) \). Another interesting observation was that the series of PA K-vi showed high gas permeability \( (e.g., P_{CO_2} = 97, P_{O_2} = 23.6) \) over the PA K-iv \( (e.g., P_{CO_2} = 83, P_{O_2} = 19.8) \) despite the presence of the additional phenyl moiety in the former. The high gas permeability values were due to the flexible ether linkage \((-O-\) between the two phenyl rings in the former. In this series of PAs, the FFV morphology was analyzed using MD simulation to understand the effect of the PA structure on the FV distribution and their gas transport properties. From Figure 17, it can be observed that the well-connected large microcavities (FV elements) of the PA L-ii was responsible for its high permeability and selectivity. The microcavities gradually become smaller and discontinuous in the following order, from PA L-vi > PA L-iv > PA L-v, i.e., with the decrease in the FFV size of the PA building blocks.

6 PAs containing Tröger’s base moiety

In recent days, polymers of intrinsic microporosity (PIMs) have drawn much attention. The PIMs have high free volume and are rapidly expanding a class of solution-processable amorphous polymer. Several types of PIMs with a bulky and kinked structure have been reported, of which Tröger’s base (TB) is one of them. TB is a bridge bicyclic diamine 2,8-dimethyl-6H,12H-5,11-methanodibenzo\([b,f][1,5\)diazocine (95). Generally, \(N\)-containing functional groups in polymer structure increase \(CO_2\)
affinity, which gives high CO$_2$ solubility in the polymer matrix. As a result, solubility and selectivity increase. The alkaline $N$ atom of a tertiary amine in TB is mainly responsible for the high affinity between CO$_2$ and polymer matrix (96). TB-based intrinsic microporosity combined with rigid PAs is expected to be a good performer in gas

Figure 15: Structure of the PAs having TPA containing pendant trityl moiety (94).

Figure 16: Free volume distributions (FVDs) of the PAs (PA I to PA V = PA K-i to PA K-v in Figure 14) (93). Copyright 2015, Reproduced with permission from American Chemical Society.
separation applications with good permeability and selectivity. In this context, Banerjee et al. has synthesized four aromatic PAs (Figure 18) with TB-derived diamine 

\[
2,8\text{-diamino-4,10-dimethyl-6H,12H-5,11-methanodibenzo[1,5]diazocine}\]

and four different aromatic acids (97). The polymers were soluble in DMAc, and membranes were fabricated through solution casting. The PAs showed thermal stability (T\text{d}0) in the range of 351–461°C, and high glass transition temperature, 283–290°C. The polymers showed high gas permeability (P\text{CO}_2 up to 109.1 Barrer and P\text{O}_2 up to 21.6 Barrer) and maintained a good selectivity. The –CH$_3$ group in TB-moiety reduces close packing between two polymer chains; thus, openness and permeability increase. Gas permeability of the membranes depends on interchain spacing (d-spacing) and FFV (Table 11). Gas permeability of this series of PAs followed the order PA M-i > PA M-ii > PA M-iii > PA M-iv, depending on their chemical structure.

In the series, PA M-i has the highest gas permeability, the presence of 6F groups in this polymer is responsible for its relatively higher permeability. TB units in the polymer backbone also improve the permeability compared to the analogous 6F group containing polymers (74,78). The presence of flexible ether linkages in the polymer chain makes PA M-ii more permeable than PA M-iii or PA M-iv. The PA M-iii showed higher permeability than PA M-iv because of the unsymmetrical structure of the acid moiety, except CO$_2$ gas. PA M-iv has the lowest FFV and consequently the lowest permeability; this is due to their high chain packing density. The ideal selectivity, CO$_2$/CH$_4$ and O$_2$/N$_2$ were in the range of 48.4–53.6 and 6.5–7.5, respectively. High selectivity for CO$_2$/CH$_4$ and O$_2$/N$_2$ gas pairs was because of the amide bonds that enhance intramolecular rigidity by densifying the polymer chains. Incorporation of TB units into the polymer backbone notably increases the selectivity with moderate permeability.

7 Future outlook

Polymeric membrane-based gas separation is an essential unit operation in many industries to recover valuable gases and facilitate pollution control and reduce capital investment, operating cost, and process safety (5–8). The success of separation of gas mixtures to their individual components from a feed lies in designing new membranes that can address the trade-off relationship between gas permeability and selectivity (3,9). The PAs are an essential class of high-performance polymers and
had shown great promise in this direction, though their utilization as gas separation membranes is not fully exploited (3). The improvement in PAs' solubility without affecting their properties is the scope of research in the field. The incorporation of different packing disrupting bulky pendant groups resulted in an increase in the PAs' solubility in selective solvents and allowed of fabrication films by solution casting route (1). It is observed that structural modifications of the PAs by incorporating different bulky groups largely affect their gas permeation behavior. The effect of structural alterations on the PAs (Tables 1–11) gas permeation properties, the selectivity values of O2/N2 and CO2/CH4 gas pairs have been presented against the O2 and CO2 permeability values, respectively (Figures 19 and 20). From the plots, it can be observed that the incorporation of the bulky groups, especially the packing disrupting tert-butyl, biphenyl fluorene, and adamantane, showed an exciting combination of gas permeability and selectivity (77,78,92,93). In O2/N2 gas pairs, the PAs having those bulky groups exceeded or closed to the Robeson present upper bound. Whereas for CO2/CH4 gas pair, the PAs having the above-described bulky groups come in close proximity to the upper bound and even touched in case of the bulky trityl and tert-butyl containing PAs. The propeller-shaped TPA core and the bulky pendant group in PAs showed a synergistic effect on both gas permeability and permselectivity (94). The observed properties make this structural variation a useful tool for designing the polymers for superior gas permeation properties. TB-based polymers are absorbing and provides both high permeability and selectivity (97). The designing of new TB-based PAs by incorporating bulky trityl and tert-butyl moieties may be beneficial in improving their separation performance. The molecular dynamic simulation can be employed forehand to predict the gas separation behaviors and understand which structural variation is paramount to obtain better separation efficiency. Also, experimental techniques like positron annihilation may enlighten better to understand the gas separation performance of the prepared membranes by calculating their free volume.

**Table 11:** FFV, Tg, and gas permeability (in Barrers) and perm-selectivity values of the TB-based PAs measured at 3.5 bar pressure and 35°C

<table>
<thead>
<tr>
<th>Polymer</th>
<th>FFVEXP</th>
<th>Tg (°C)</th>
<th>P CO2</th>
<th>P O2</th>
<th>α (CO2/CH4)</th>
<th>α (O2/N2)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>PA M-i</td>
<td>0.245</td>
<td>283</td>
<td>109.1</td>
<td>21.6</td>
<td>50.3</td>
<td>7.52</td>
<td>(97)</td>
</tr>
<tr>
<td>PA M-ii</td>
<td>0.229</td>
<td>284</td>
<td>92.2</td>
<td>18.3</td>
<td>46.8</td>
<td>7.20</td>
<td></td>
</tr>
<tr>
<td>PA M-iii</td>
<td>0.168</td>
<td>285</td>
<td>84.7</td>
<td>16.8</td>
<td>48.4</td>
<td>7.21</td>
<td></td>
</tr>
<tr>
<td>PA M-iv</td>
<td>0.166</td>
<td>290</td>
<td>86.4</td>
<td>14.1</td>
<td>53.7</td>
<td>6.52</td>
<td></td>
</tr>
</tbody>
</table>

**Figure 19:** Robeson plot of permeability/selectivity trade-off for O2 and N2 gas pair. Values are taken from Tables 1–11.

**Figure 20:** Robeson plot of permeability/selectivity trade-off for CO2 and CH4 gas pair. Values are taken from Tables 1–11.

8 Summary

In summary, the PA membranes are making steady progress for being used in membrane-based gas separation applications. But still, some issues should be considered for the further development of the PAs as membrane materials for gas separation applications. First of all, there remains the trade-off between gas permeability and selectivity, as expressed in Robeson’s upper bound. Development of PAs with proper selection of bulky
pendant groups like tert-butyl, combined with a propeller-shaped TPA structure, can be a way forward to achieve higher permeability with higher selectivity. However, mechanical properties tend to decrease when bulky groups are incorporated in the PAs. The bulky groups reduce the H-bonding between the polymer chains and disrupt chain packing, affecting the membrane suitability at high-pressure, high-temperature gas separation applications. Another approach should also be considered, mixing the PAs with the inorganic filler particles for the preparation of the mixed matrix membranes. This approach will allow the opportunity to effectively use the strength, stability, and permeation efficiency of the PAs with the better selective properties of the inorganic fillers. Also, cardo groups with different functionalities in the PAs may help to achieve intended gas separation efficiency. Physical aging and plasticization of the polymer membranes are the significant issues for successful membrane-based separation for commercial applications (98). Hence, the gas-separation studies with aged PA membranes also need to be investigated to qualify these classes or polymers for an industrial separation process.

**Research funding:** This work was supported by Wacker Metroarck Chemicals Pvt. Ltd. through a sponsored research agreement (Project No. WACKER-IIT Kgp 2020/03).

**Authors contribution:** Debaditya Bera: writing – original draft, diagram, and figures; Rimpa Chatterjee: writing – original draft, data collection and analysis; Susnata Banerjee: conceptualization, writing – review and editing, visualization, project administration, funding acquisition.

**Conflict of interest:** One of the authors (Susanta Banerjee) is a member of the Editorial Advisory Board of e-Polymers.

**References**


18. Sen SK, Banerjee S. Gas transport properties of fluorinated poly (ether imide) films containing phthalimidine moiety in the...


(46) Maji S, Banerjee S. Synthesis and characterization of new meta connecting semifluorinated poly (ether amide) s and their


