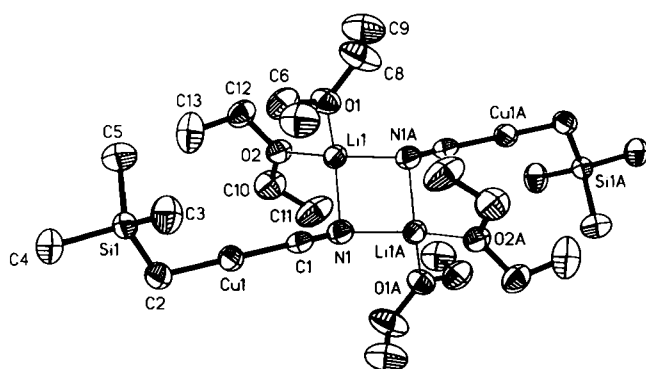


Crystal structure of bis[bis(diethylether)lithium (trimethylsilyl)methyl-cyanocuprate], [(C₄H₁₀O)₂Li(C₃H₉)SiCH₂CuCN]₂

F. Bosold, M. Marsch, K. Harms* and G. Boche

Philipps Universität, Fachbereich Chemie, Hans-Meerwein-Str., D-35032 Marburg, Germany

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**Table 1.** Data collection and handling.

Crystal:	colourless prism, size 0.4 × 0.4 × 1.0 mm
Wavelength:	Mo K α radiation (0.71073 Å)
μ :	11.87 cm ⁻¹
Diffractometer, scan mode:	Stoe IPDS, $\Delta\phi = 1^\circ$
2 θ _{max} :	51.78°
$N(hkl)$ _{measured} , $N(hkl)$ _{unique} :	14200, 3585
Criterion for I_{obs} , $N(hkl)$ _{gt} :	$I_{obs} > 2 \sigma(I_{obs})$, 2601
$N(param)$ _{refined} :	179
Programs:	SHELXS-97 [4], SHELXL-97 [5], DIFABS [6]

Abstract

C₂₆H₆₂Cu₂Li₂N₂O₄Si₂, monoclinic, $P12_1/c1$ (No. 14),
 $a = 10.0080(8)$ Å, $b = 11.1636(6)$ Å, $c = 18.029(2)$ Å,
 $\beta = 106.091(9)^\circ$, $V = 1935.4$ Å³, $Z = 2$, $R_{gt}(F) = 0.031$,
 $wR_{ref}(F^2) = 0.074$, $T = 193$ K.

Source of material

The synthesis was done according to [1].

Discussion

The first X-ray crystal structure of a lower order cyano cuprate [1], [tBuCuC(CN)Li(OEt₂)₂]_∞, has been published only recently [2]. The most significant different between this structure and the structure published here is the absence of a Cu^I(d¹⁰)–Cu^I(d¹⁰) bonding interaction in the latter one (see also [3]).

Table 2. Atomic coordinates and displacement parameters (in Å²).

Atom	Site	x	y	z	U_{iso}
H(2A)	4e	0.8314	1.0246	0.3728	0.057
H(2B)	4e	0.9372	0.9255	0.3593	0.057
H(3A)	4e	1.0942	1.0719	0.2375	0.092
H(3B)	4e	1.2001	1.0094	0.3103	0.092
H(3C)	4e	1.2134	1.1495	0.2949	0.092
H(4A)	4e	1.0807	1.1445	0.4884	0.093
H(4B)	4e	1.2031	1.1995	0.4577	0.093
H(4C)	4e	1.1939	1.0579	0.4690	0.093
H(5A)	4e	0.8999	1.2685	0.2684	0.094
H(5B)	4e	1.0252	1.3284	0.3328	0.094
H(5C)	4e	0.8866	1.2884	0.3540	0.094
H(6A)	4e	0.7583	1.2215	0.1302	0.078
H(6B)	4e	0.8021	1.3310	0.0849	0.078
H(7A)	4e	0.9793	1.1928	0.1196	0.112
H(7B)	4e	0.9238	1.2012	0.0275	0.112
H(7C)	4e	0.8788	1.0909	0.0720	0.112
H(8A)	4e	0.7425	1.3342	-0.0427	0.101
H(8B)	4e	0.6599	1.2237	-0.0908	0.101
H(9A)	4e	0.5369	1.3850	-0.1334	0.138
H(9B)	4e	0.5471	1.4300	-0.0477	0.138
H(9C)	4e	0.4553	1.3149	-0.0819	0.138
H(10A)	4e	0.2460	1.2087	0.1350	0.081
H(10B)	4e	0.3594	1.1182	0.1857	0.081
H(11A)	4e	0.1789	1.0111	0.1064	0.125
H(11B)	4e	0.3178	0.9874	0.0810	0.125
H(11C)	4e	0.2018	1.0771	0.0320	0.125
H(12A)	4e	0.3743	1.3664	0.1244	0.067
H(12B)	4e	0.5189	1.3541	0.1035	0.067
H(13A)	4e	0.5710	1.3983	0.2322	0.127
H(13B)	4e	0.6132	1.2610	0.2267	0.127
H(13C)	4e	0.4717	1.2940	0.2470	0.127

* Correspondence author (e-mail: harms@chemie.uni-marburg.de)

Table 3. Atomic coordinates and displacement parameters (in Å²).

Atom	Site	x	y	z	U ₁₁	U ₂₂	U ₃₃	U ₁₂	U ₁₃	U ₂₃
Cu(1)	4e	0.77836(3)	0.97960(3)	0.23662(2)	0.0427(2)	0.0424(2)	0.0404(2)	-0.0061(1)	0.0080(1)	0.0001(1)
Si(1)	4e	1.02747(6)	1.11692(6)	0.35167(3)	0.0400(3)	0.0413(4)	0.0355(3)	-0.0004(3)	0.0110(2)	-0.0010(3)
C(1)	4e	0.6668(2)	0.9570(2)	0.1356(1)	0.041(1)	0.029(2)	0.049(1)	-0.0030(9)	0.012(1)	0.002(1)
O(1)	4e	0.6555(2)	1.2268(2)	0.01875(9)	0.0445(9)	0.056(1)	0.0511(9)	-0.0084(8)	0.0168(7)	0.0100(8)
N(1)	4e	0.5990(2)	0.9449(2)	0.0736(1)	0.042(1)	0.038(1)	0.047(1)	-0.0024(8)	0.0091(9)	0.0015(8)
Li(1)	4e	0.5097(4)	1.1194(3)	0.0313(2)	0.042(2)	0.036(2)	0.041(2)	-0.003(2)	0.015(2)	-0.003(2)
O(2)	4e	0.4200(2)	1.1994(2)	0.10204(9)	0.0508(9)	0.039(1)	0.0540(9)	-0.0011(7)	0.0280(8)	-0.0025(7)
C(2)	4e	0.8922(2)	1.0027(2)	0.3400(1)	0.052(1)	0.049(2)	0.043(1)	-0.007(1)	0.012(1)	0.003(1)
C(3)	4e	1.1482(3)	1.0829(3)	0.2914(2)	0.054(2)	0.077(2)	0.057(2)	-0.004(1)	0.023(1)	-0.011(1)
C(4)	4e	1.1398(3)	1.1314(3)	0.4539(1)	0.067(2)	0.066(2)	0.048(1)	-0.007(1)	0.005(1)	-0.009(1)
C(5)	4e	0.9507(3)	1.2686(3)	0.3233(2)	0.064(2)	0.049(2)	0.076(2)	0.003(1)	0.019(1)	0.007(1)
C(6)	4e	0.7787(3)	1.2446(3)	0.0815(2)	0.058(2)	0.056(2)	0.079(2)	-0.016(1)	0.016(1)	-0.011(2)
C(7)	4e	0.9002(3)	1.1767(3)	0.0746(2)	0.049(2)	0.079(3)	0.092(2)	-0.006(2)	0.010(2)	0.011(2)
C(8)	4e	0.6571(4)	1.2852(3)	-0.0517(2)	0.092(2)	0.089(3)	0.087(2)	0.015(2)	0.048(2)	0.042(2)
C(9)	4e	0.5402(4)	1.3594(4)	-0.0809(2)	0.087(2)	0.077(3)	0.110(3)	0.012(2)	0.022(2)	0.036(2)
C(10)	4e	0.3160(3)	1.1473(3)	0.1326(2)	0.069(2)	0.067(2)	0.084(2)	-0.010(2)	0.049(2)	-0.003(2)
C(11)	4e	0.2480(3)	1.0475(3)	0.0840(2)	0.085(2)	0.055(2)	0.135(3)	-0.019(2)	0.074(2)	-0.020(2)
C(12)	4e	0.4592(3)	1.3167(2)	0.1326(2)	0.062(2)	0.044(2)	0.068(2)	0.003(1)	0.029(1)	-0.010(1)
C(13)	4e	0.5354(3)	1.3176(4)	0.2169(2)	0.074(2)	0.110(3)	0.071(2)	-0.006(2)	0.021(2)	-0.030(2)

References

- Boche, G.; Bosold, F.; Marsch, M.; Harms, K.: Die Kristallstrukturen eines Lower-order- und eines Higher-order-Cyanocuprates: [tBuCu(CN)Li(OEt₂)₂]_∞ und [tBuCuBu{Li(thf)(pmdeta)₂CN}]. *Angew. Chem.* **110** (1998) 1779-1781; *Ang. Chem. Int. Ed. Engl.* **37** (1998) 1684-1686.
- Krause, N.: Highlight: Neues zur Struktur und Reaktivität von Cyanocupraten - das Ende einer alten Kontroverse. *Angew. Chem.* **111** (1999) 83-85; *Ang. Chem. Int. Ed. Engl.* **38** (1999) 79-81.
- John, M.; Auel, C.; Behrens, Chr.; Marsch, M.; Harms, K.; Bosold, F.; Gschwind, R. M.; Rajamohanam, P. R.; Boche, G.: The Relation between Ion Pair Structures and Reactivities of Lithium Cuprates. *Chem. Eur. J.* **6** (2000) 3060-3068.
- Sheldrick, G.M.: SHELXS-97. Program for Crystal Structure Solution. University of Göttingen, Germany 1997.
- Sheldrick, G.M.: SHELXL-97. Program for Crystal Structure Refinement. University of Göttingen, Germany 1997.
- Walker, N.; Stuart, D.: An Empirical Method for Correcting Diffractometer Data for Absorption Effects. *Acta Crystallogr. A* **39** (1983) 158-166.