Synthesis of the Bifunctional Ligand \(N,N'[1,4-
\text{Phenylenebis(iminocarbonyl)}]dil(\text{L-phenylalanine})\) and Crystal Structure of Its Supramolecular Coordination Polymer with Lead(II)

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The new linker-type ligand 1, featuring two L-phenylalanine-terminated urea moieties attached to the para positions of a central phenylene unit has been synthesized. Its coordination polymer formed with lead(II) and containing DMF and water as solvent molecules has been prepared and structurally studied by X-ray crystallography.

Key words: \(N,N'[1,4-
\text{Phenylenebis(iminocarbonyl)}]dil(\text{L-phenylalanine})\), Lead(II) Complex, Synthesis, Crystal Structure, Supramolecular Coordination Polymer

Introduction

Coordination polymers, also known as metal-organic framework materials, are of high current interest [1 – 3] due to their potential uses in different fields of application including catalysis [4 – 6], selective gas adsorption [7 – 9], chromatography [10 – 12], magnetism [2, 13 – 15], and non-linear optical behavior [16 – 18]. Materials of this type owe their properties to a clever choice of the building components, that is the organic linker molecule and the coordinating metal ion. Nevertheless, this is not an easy point because of interpenetration and connectivity problems interfering with the rational design and synthesis of an intended coordination network architecture [2, 19]. Among the great variety of organic linker molecules that have been tested [1, 2, 20], the molecules possessing carboxylic acid functions as coordinating end groups were shown to be particularly efficient. A great many of metal ions have been used for coordination to carboxylic linker molecules including Zn(II), Cu(II), Ni(II) and others [1, 2, 20], resulting in the formation of specific points of coordinative intersection referred to as SBU's in this context.

Owing to the 6\(\sigma^2\) lone pair of divalent lead, which can cause distortion of the coordination sphere [21 – 23] frequently discussed in considering its stereochemical activity [24 – 26], Pb(II) plays an important role in coordination chemistry since it can lead to either holodirected or hemidirected coordination [27 – 29] dependent on the nature of the ligand such as containing soft or hard donor atoms or attractive versus repulsive interactions among ligands [30]. Hence, Pb(II) complexed to carboxylic ligands, giving rise to both chelating and bridging coordination modes, offers suitable requirements for the construction of polymeric coordination networks [31, 32]. This has been documented in a recent very comprehensive review article showing numerous examples of respective complexes [33]. However, the corresponding complexes of Pb(II) with amino acid ligands [34 – 40], including phenylalanine [40 – 42], are still not so common, and related complexes formed of a ditopic linker.
molecule based on an amino acid derivative, as far as we know, are not in the literature.

Here we report the coordination polymer of such a new bifunctional amino acid-derived linker molecule featuring two 1-phenylalanine-terminated urea moieties attached to the para positions of a central phenylene unit with Pb(II). We describe the synthesis of the linker compound as well as the preparation of the coordination polymer and in particular discuss the X-ray crystal structure of the latter species paying – aside from the coordination mode of the metal ion – also attention to the supramolecular behavior of the urea moieties to control the packing.

Experimental Section

General

The melting points were determined with a Melting Point B-540 (Buchi, Switzerland) and are uncorrected. The elemental analyses were determined on a Heraeus CHN rapid analyzer. The IR spectra were obtained with a Nicolet 510 FT-IR spectrometer. 1H and 13C NMR spectra were recorded using a Bruker Avance Ultrasound 500 (500 MHz) instrument. The chemical shifts (δ) are reported as ppm relative to SiMe4. The mass spectrum (ESI) was recorded with a Varian LCMS-320 spectrometer. Organic solvents were purified by standard procedures. Starting compounds p-phenylene diisocyanate, L-phenylalanine and lead(II) nitrate were purchased from commercial suppliers.

Synthesis of N,N′-[1,4-phenylenebis(iminocarbonyl)]di(L-phenylalanine), LH2 (I)

L-Phenylalanine (1.65 g, 10 mmol) and NaOH (0.41 g, 10 mmol) were dissolved in water (12.5 mL) and NaOAc (0.79 g, 5 mmol) was added in toluene (8.5 mL) was added under stirring at 0°C. After completion of the reaction (4.5 h) the phases were separated. The aqueous phase was acidified with 3 N hydrochloric acid and the precipitate collected. Recrystallization from methanol of the solvent, the colorless crystals were collected and dried to yield 74% of 1a; m. p. 228–230°C (dec.). – IR (KBr): νmax = 1651 (C=O, urea), 1559 (COO−), 1502 (C=C), 1401 (COO−), 827, 701 (C-H, Ar). – C26H26N2O8Pb: calcd. C 59.79, H 5.85, N 10.83; found C 59.79, H 5.85, N 10.83.

Preparation of PbLH2·H2O·DMF (1a)

Equimolar amounts of 1 and lead(II) nitrate were dissolved in dimethylformamide. After isothermal evaporation of the solvent, the colorless crystals were collected and dried to give 74% of 1a; m. p. 228–230°C (dec.). – IR (KBr): νmax = 1651 (C=O, urea), 1559 (COO−), 1502 (C=C), 1401 (COO−), 827, 701 (C-H, Ar). – C26H26N2O8Pb: calcd. C 44.27, H 4.23, N 8.90; found C 44.54, H 4.52, N 9.52.

X-Ray structure determination

Intensity data were collected on a Kappa APEX II diffractometer (Bruker AXS) using ω- and φ-scans. The collected data were corrected for Lorentz and polarization effects. The structure was solved by Direct Methods (SHELXS-97 [43]) and refined by full-matrix least-squares calculations on F2 (SHELXL-97 [44]). Empirical absorption correction based on multi-scans was applied by using the program SADABS [45]. The non-hydrogen positions were refined together with their anisotropic displacement parameters; the H atoms were treated isotropically. The hydrogen atoms of the water molecule could not be located from a difference electron density map and thus are not included in the structure model. All other H positions were held riding on the respective parent C and N atoms during the subsequent calculations.

CCDC 885573 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Results and Discussion

The ligand LH2 (Fig. 1), was synthesized by addition of L-phenylalanine to p-phenylene diisocyanate following a related literature procedure [46]. The com-

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Fig. 1. Compounds studied in this article.
with the ligand and one molecule of DMF coordinated to the Pb(II) ion (Fig. 2). The proximity of the water oxygen to one of the carboxylate oxygen atoms [d(O···O) 2.919(9) Å] suggests molecular association via O–H···O bonding. The conformation of the ligand molecule (L) can be expressed by a set of torsion and dihedral angles (Table 2), the latter describing the orientation of the planar molecular elements to each other. The terminal aromatic rings of L adopt approx-
The Pb(II) center adopts a strongly distorted square-pyramidal coordination environment with the oxygen atom of DMF and the 6s lone pair on lead occupying the axial positions (Fig. 3). Within the coordination polyhedron, the metal ion is located at a distance of ca. 0.75 Å from the basal plane which is defined by non-equivalent carboxylate oxygen atoms [O(3)–O(6)] of four different ligand molecules. The Pb–O distances range from 2.373(6) to 2.752(7) Å, the O–Pb–O angles are between 67.1(2) and 142.6(2)°. Two additional weak interactions with Pb–O distances of 2.824(7) and 2.885(7) Å complete the coordination environment of the metal ion. As displayed in Fig. 3(b), the bonds to the ligand molecules are asymmetrically distributed around the metal center and directed to only imitate coplanarity [2.9(6)°], but are inclined at angles of 79.0(3) and 81.4(3)° with respect to the plane of the central arene ring. The dihedral angles formed between the planes through the urea fragments and the phenylene ring are 49.5(3)°.

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\begin{align*}
\text{Dihedral angles (deg)}^a & \\
\text{mpln(A)-mpln(B)} & 79.0(3) \quad \text{mpln(A)-mpln(D)} & 74.3(3) \\
\text{mpln(A)-mpln(C)} & 2.9(6) \quad \text{mpln(C)-mpln(E)} & 77.4(3) \\
\text{mpln(B)-mpln(C)} & 81.4(3) \quad \text{mpln(B)-mpln(F)} & 49.5(3) \\
\text{Bond lengths (Å)} & \\
\text{Pb(1)–O(3)}^i & 2.373(6) \quad \text{Pb(1)–O(1A)} & 2.454(6) \\
\text{Pb(1)–O(4)}^i & 2.387(6) \quad \text{Pb(1)–O(3)} & 2.824(6) \\
\text{Bond angles (deg)} & \\
\text{Pb(1)–O(3)} & 2.373(6) \quad \text{Pb(1)–O(1A)} & 2.454(6) \\
\text{Pb(1)–O(4)} & 2.387(6) \quad \text{Pb(1)–O(3)} & 2.824(6) \\
\end{align*}
\]

\[\text{Symmetry codes: (i) } 1 - x, y, z; \text{(ii) } x, 1 + y, z; \text{(iii) } 1 - x, 1 + y, -1 + z; \text{(iv) } x, 1 + y, -1 + z.\]

\[\begin{align*}
\text{D–H···A} & \quad \text{Symmetry} & \text{D–H} & \text{D–A} & \text{H···A} & \text{Angles (deg)} \quad \text{D–H···A} \\
\text{N(1)–H1N1} & -1 + x, y, z & 0.86 & 2.933(9) & 2.17 & 147 \\
\text{N(2)–H1N2} & -1 + x, y, z & 0.86 & 2.903(9) & 2.11 & 153 \\
\text{N(3)–H1N3} & -1 + x, y, z & 0.86 & 2.871(11) & 2.06 & 157 \\
\text{N(4)–H1N4} & -1 + x, y, z & 0.86 & 2.985(9) & 2.26 & 142 \\
\text{C(1A)–H1A} & x, 1 + y, -1 + z & 0.93 & 3.008(12) & 2.42 & 122 \\
\text{C(8)–H8} & -1 + x, y, z & 0.98 & 3.126(10) & 2.35 & 135 \\
\text{C(3A)–H3A} & x, y, z & 0.96 & 2.777(15) & 2.41 & 103 \\
\text{C(3A)–H3A} & -1 + x, y, z & 0.96 & 3.490(14) & 2.54 & 172 \\
\text{C(17)–H17} & x, y, z & 0.98 & 2.740(11) & 2.41 & 104 \\
\text{C(17)–H17} & -1 + x, y, z & 0.98 & 3.190(11) & 2.32 & 148 \\
\text{C(15)–H15} & x, 1 + y, z & 0.93 & 3.603(12) & 2.81 & 144 \\
\text{C(12)–H12} & x, -1 + y, z & 0.93 & 3.519(12) & 2.75 & 141 \\
\end{align*}\]

\[\begin{align*}
\text{Symmetry codes: (i) } x, y, z; \text{(ii) } 1 - x, 1 + y, z; \text{(iii) } 1 - x, 1 + y, -1 + z; \text{(iv) } x, 1 + y, -1 + z.\]

\[\begin{align*}
\text{mpln(A)-mpln(B)} & 79.0(3) \quad \text{mpln(A)-mpln(D)} & 74.3(3) \\
\text{mpln(A)-mpln(C)} & 2.9(6) \quad \text{mpln(C)-mpln(E)} & 77.4(3) \\
\text{mpln(B)-mpln(C)} & 81.4(3) \quad \text{mpln(B)-mpln(F)} & 49.5(3) \\
\end{align*}\]
Fig. 4. Packing of the coordination polymer viewed down the crystallographic a axis. Oxygen atoms are displayed as grey, nitrogen atoms as hatched and metal atoms as cross-hatched circles. Broken lines represent hydrogen bond interactions.

a part of the coordination sphere. This hemidirected coordination geometry which is induced by the lone pair of electrons [24–26] is found in many Pb(II) complexes [27–29, 33].

According to the given coordination mode, the crystal structure is composed of two-dimensional supramolecular networks extending parallel to the crystallographic[022] plane. In this arrangement, the urea fragments form the well-known α-networks [47] assembled via N–H···O=C hydrogen bonds [d(N···O) 2.871(11)–2.985(9) Å] with the oxygen atoms acting as bifurcated acceptors (Fig. 3). This pattern of hydrogen bonding follows the graph set C(4)[R12(6)] [48] created by both a six-membered hydrogen bonded ring and a C(4) chain. It should be noted at this point that the formation of supramolecular tapes of this kind is observed in the solid-phase structures of a variety of N,N'-disubstituted derivatives of urea [48–51], which in many cases have been used as a control element for crystal engineering and template-directed supramolecular synthesis [52–54].

In the present structure, the two sets of α-networks have opposite directions and are assisted by π···π stacking [55] of the neighboring aromatic rings. The center–center distance of 4.655 Å between interacting rings corresponds with the length of the a axis of the unit cell. A view of the packing along this axis reveals a channel-like cavity structure which is partially occupied by solvent molecules (Fig. 4).

Moreover, the crystal contains potential solvent-accessible voids of 67.6 Å³ per unit cell, which represents 8.4% of the total cell volume. The 2D networks are associated by weak C–H···π arene interactions [56] [C(15)–H(15)···C(21) 2.81 Å, 144°; C(12)–H(12)···C(3) 2.75 Å, 141°].

Conclusions

A new ditopic linker molecule 1, synthesized from L-phenylalanine and p-phenylene diisocyanate, has been found to form a crystalline coordination polymer with Pb(II) containing a hydrogen-bonded supramolecular network structure and getting additional support from π···π stacking interactions. Induced by its 6s² lone pair, the Pb(II) centers are asymmetrically seven-coordinated including six oxygen atoms of carboxylate groups and the oxygen atom of a DMF solvent molecule, thus giving rise to a hemidirected coordination geometry, rather common in the complexation of Pb(II) [33]. Moreover, the urea subunits show their typical behavior as synthons [57], which is the hydrogen bonded assembly of six-membered rings [58], while π···π stacking interactions are due to the presence of the central phenylene rings of the linker molecule. This cooperative binding situation, structurally programmed in the linker molecule, finally leads to the formation of the particular two-dimensional coordination polymer which may be seen as a hybrid constellation between coordination polymer and organic framework structure [32, 59–61]. Beyond that, caused by steric effects of the terminal arene groups, channel-like lattice voids are created which are partially filled with solvent molecules, giving the impression of a kind of clathrate formation [62]. As much as 8.4% of the total crystal volume represents empty space so that the crystal also features a porous cavity structure [63].

Based on this property, the present ligand design may arouse interest in the community of re-
searchers dealing with crystal engineering and materials chemistry. Besides, the enantiopure structure of the molecule could open a further direction of development in this field regarding its optical behavior.

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